Fundamentals of organic chemistry solomons pdf

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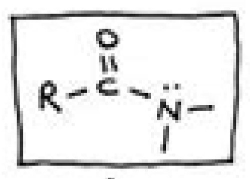
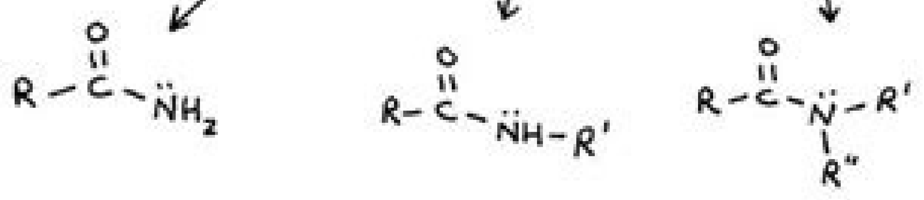
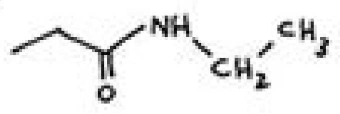
④ Amides ⇒  ≡ "alkanamide"

a). Straight-chained amides.

↳ can be 1° amide (primary), 2° amide (secondary), or 3° amide (tertiary)

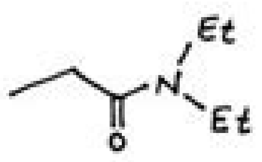


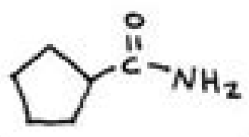
ex: $CH_3CH_2CH_2-\overset{O}{\overset{\|}{C}}-NH_2$ → butanamide (1° amide)

ex:  → N-ethylpropanamide (2° amide)

↳ "propanamide part"

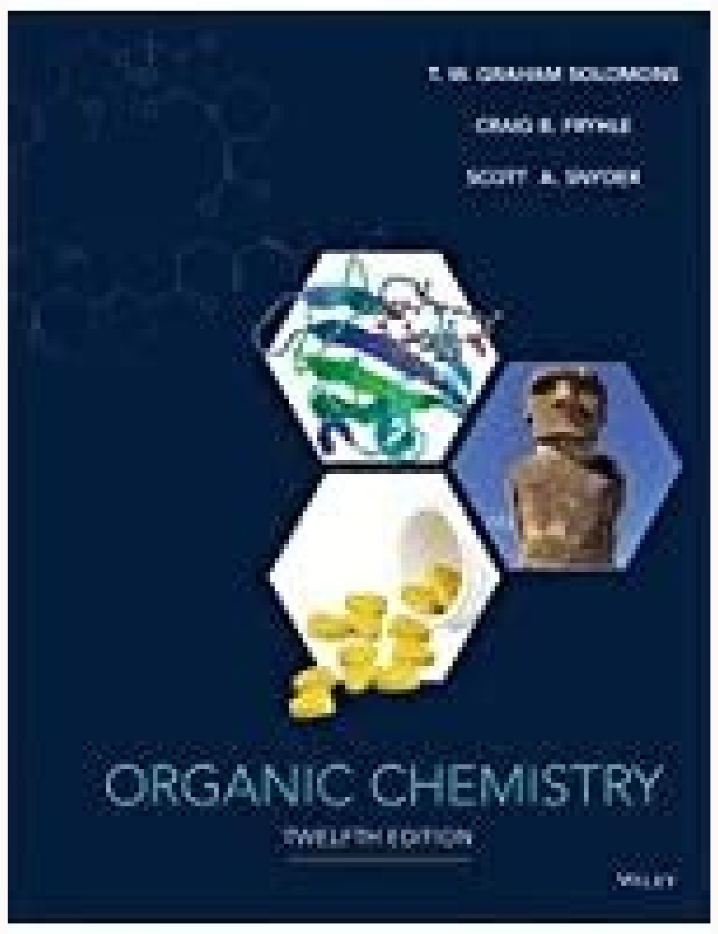
↳ denotes that ethyl group is "off the N".

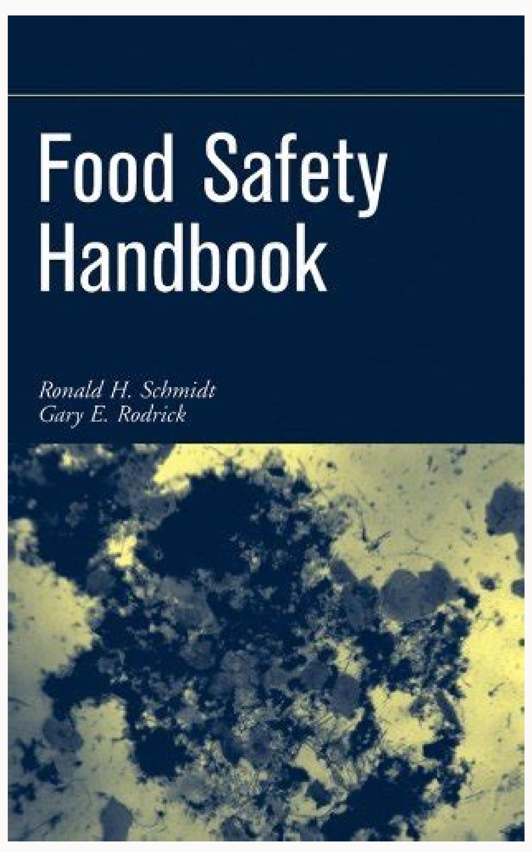
ex:  → N,N-diethylpropanamide (3° amide)

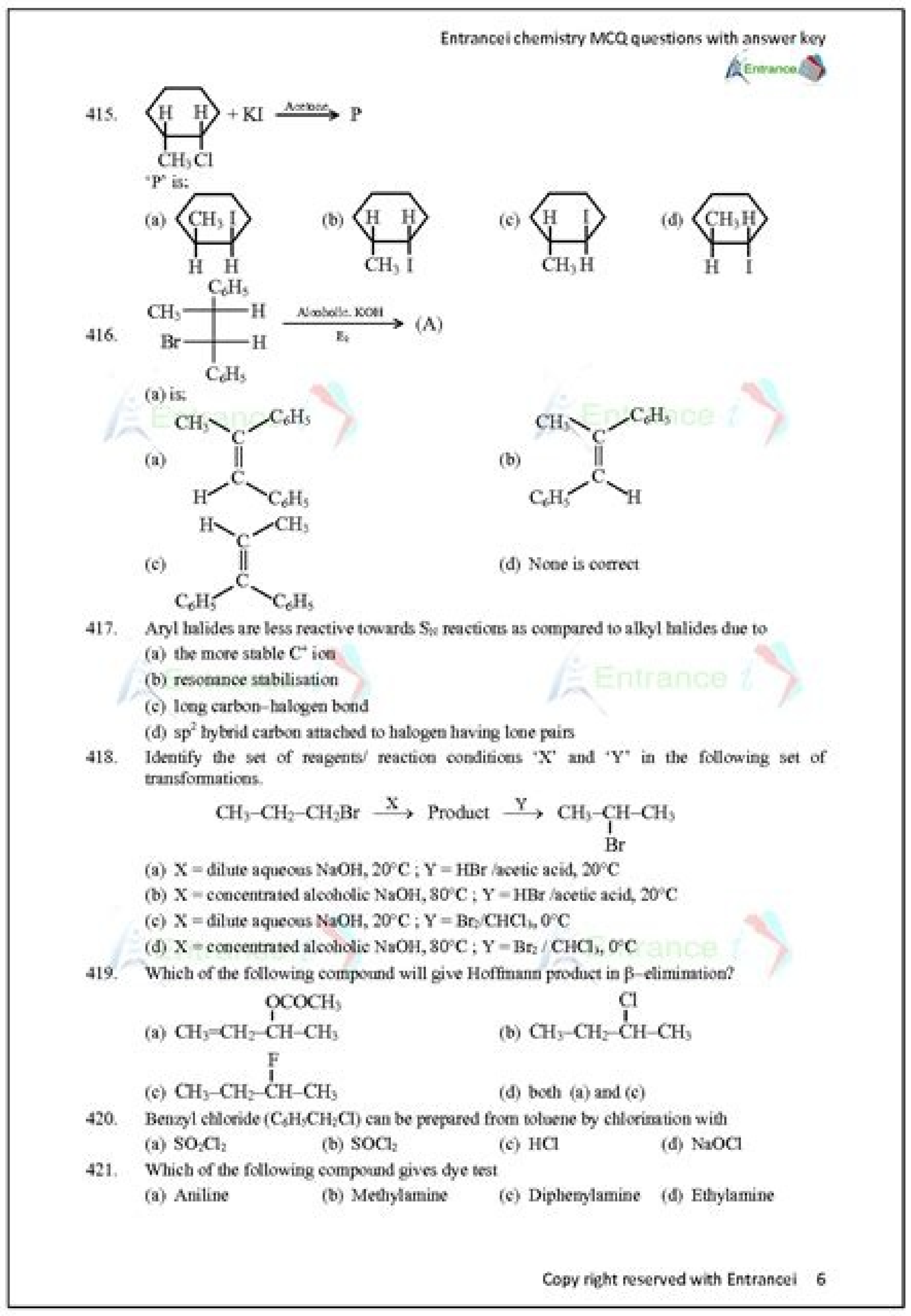
ex:  → cyclopentanecarboxamide

↳ NOT a "cyclic amide" -- see next page...

	Class	Suffix Name	Prefix Name
↑ increasing priority	Carboxylic acid	-oic acid	Carboxy
	Ester	-oate	Alkoxycarbonyl
	Amide	-amide	Amido
	Nitrile	-nitrile	Cyano
	Aldehyde	-al	Formyl (—CH=O)
	Ketone	-one	Oxo (=O)
	Alcohol	-ol	Hydroxy
	Amine	-amine	Amino
	Alkene	-ene	Alkenyl
	Alkyne	-yne	Alkynyl
	Alkane	-ane	Alkyl
	Ether	—	Alkoxy
	Alkyl halide	—	Halo






Entrancei chemistry MCQ questions with answer key

415. + KI $\xrightarrow{\text{Acetone}}$ P

'P' is:

(a) (b) (c) (d)

416. $\xrightarrow[E_2]{\text{Alcoholic KOH}}$ (A)

(a) is:

(a) (b)

(c) (d) None is correct

417. Aryl halides are less reactive towards S_N reactions as compared to alkyl halides due to

(a) the more stable C^+ ion

(b) resonance stabilisation

(c) long carbon–halogen bond

(d) sp^2 hybrid carbon attached to halogen having lone pairs

418. Identify the set of reagents/ reaction conditions 'X' and 'Y' in the following set of transformations.

$$CH_3{-}CH_2{-}CH_2Br \xrightarrow{X} \text{Product} \xrightarrow{Y} CH_3{-}\underset{\substack{|\\ Br}}{CH}{-}CH_3$$

(a) X = dilute aqueous NaOH, 20°C ; Y = HBr /acetic acid, 20°C

(b) X = concentrated alcoholic NaOH, 80°C ; Y = HBr /acetic acid, 20°C

(c) X = dilute aqueous NaOH, 20°C ; Y = $Br_2/CHCl_3$, 0°C

(d) X = concentrated alcoholic NaOH, 80°C ; Y = Br_2 / $CHCl_3$, 0°C

419. Which of the following compound will give Hoffmann product in β–elimination?

(a) $CH_3{=}CH_2{-}\overset{\substack{OCOCH_3\\ |}}{CH}{-}CH_3$ (b) $CH_3{-}CH_2{-}\overset{\substack{Cl\\ |}}{CH}{-}CH_3$

(c) $CH_3{-}CH_2{-}\overset{\substack{F\\ |}}{CH}{-}CH_3$ (d) both (a) and (c)

420. Benzyl chloride ($C_6H_5CH_2Cl$) can be prepared from toluene by chlorination with

(a) SO_2Cl_2 (b) $SOCl_2$ (c) HCl (d) NaOCl

421. Which of the following compound gives dye test

(a) Aniline (b) Methylamine (c) Diphenylamine (d) Ethylamine

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You're Reading a Free Preview Page 632 is not shown in this preview. You're Reading a Free Preview Pages 738 to 791 are not shown in this preview. You're Reading a Free Preview Pages 856 to 1144 are not shown in this preview. You're Reading a Free Preview Pages 1209 to 1292 are not shown in this preview. You're Reading a Free Preview Pages 1357 to 1408 are not shown in this preview. You're Reading a Free Preview Pages 1473 to 1594 are not shown in this preview. You're Reading a Free Preview Pages 1661 to 1677 are not shown in this preview. You're Reading a Free Preview Pages 1703 to 1739 are not shown in this preview. You're Reading a Free Preview Pages 1765 to 1870 are not shown in this preview. You're Reading a Free Preview Pages 1896 to 1978 are not shown in this preview. You're Reading a Free Preview Pages 2004 to 2060 are not shown in this preview. Capacity of a substance to dissolve in a solvent in a homogeneous way This article is about a chemical property. For the algebraic concept, see solvable group. For other uses, see solution (disambiguation). Example for a dissolved solid (left) Formation of crystals in a 4.2 M ammonium sulfate solution. The solution was initially prepared at 20 °C and then stored for 2 days at 4 °C. In chemistry, solubility is the ability of a substance, the solute, to form a solution with another substance, the solvent. Insolubility is the opposite property, the inability of the solute to form such a solution. The extent of the solubility of a substance in a specific solvent is generally measured as the concentration of the solute in a saturated solution, one in which no more solute can be dissolved.[1] At this point, the two substances are said to be at the solubility equilibrium. For some solutes and solvents there may be no such limit, in which case the two substances are said to be "miscible in all proportions" (or just "miscible").[2] The solute can be a solid, a liquid, or a gas, while the solvent is usually solid or liquid. Both may be pure substances, or may themselves be solutions. Gases are always miscible in all proportions, except in very extreme situations,[3] and a solid or liquid can be "dissolved" in a gas only by passing into the gaseous state first. The solubility mainly depends on the composition of solute and solvent (including their pH and the presence of other dissolved substances) as well as on temperature and pressure. The dependency can often be explained in terms of interactions between the particles (atoms, molecules, or ions) of the two substances, and of thermodynamic concepts such as enthalpy and entropy. Under certain conditions, the concentration of the solute can exceed its usual solubility limit. The result is a supersaturated solution, which is metastable and will rapidly exclude the excess solute if a suitable nucleation site appears.[4] The concept of solubility does not apply when there is an irreversible chemical reaction between the two substances, such as the reaction of calcium hydroxide with hydrochloric acid; even though one might say, informally, that one "dissolved" the other. The solubility is also not the same as the rate of solution, which is how fast a solid solute dissolves in a liquid solvent. This property depends on many other variables, such as the physical form of the two substances and the manner and intensity of mixing. The concept and measure of solubility are extremely important in many sciences besides chemistry, such as geology, biology, physics, and oceanography, as well as in engineering, medicine, agriculture, and even in non-technical activities like painting, cleaning, cooking, and brewing. Most chemical reactions of scientific, industrial, or practical interest only happen after the reagents have been dissolved in a suitable solvent. Water is by far the most common such solvent. The term "soluble" is sometimes used for materials that can form colloidal suspensions of very fine solid particles in a liquid.[5] The quantitative solubility of such substances is generally not well-defined, however. Quantification of solubility The solubility of a specific solute in a specific solvent is generally expressed as the concentration of a saturated solution of the two.[1] Any of the several ways of expressing concentration of solutions can be used, such as the mass, volume, or amount in moles of the solute for a specific mass, volume, or mole amount of the solvent or of the solution. Per quantity of solvent In particular, chemical handbooks will often express the solubility of a substance in a liquid as grams of solute per decilitre (100 mL) of solvent (g/dL); or, less commonly, as grams per litre (g/L). The quantity of solvent can instead be expressed in mass, as in g/100g" or g/kg. The number may be expressed as a percentage in this case, and the abbreviation "w/w" may be used to indicate "weight per weight".[6] (The values in g/L and g/kg are practically the same for water, but not for other solvents.) Alternatively, the quantity of solute can be expressed in moles instead of mass; if the quantity of solvent is given in kilograms, the value is the molality of the solution (mol/kg). Per quantity of solution The solubility of a substance in a liquid may also be expressed as the quantity of solute per quantity of solution, rather than of solvent. For example, following the common practice in titration, it may be expressed as moles of solute per litre of solution (mol/L), the molarity of the latter. In more specialized contexts the solubility may be given by the mole fraction (moles of solute per total moles of solute plus solvent) or by the mass fraction at equilibrium (mass of solute per mass of solute plus solvent), both adimensional numbers between 0 and 1 which may be expressed as percentages. Liquid and gaseous solutes For solutions of liquids or gases in liquids, the quantities of both substances may be given volume rather than mass or mole amount; such as litre of solute per litre of solvent, or litre of solute per litre of solution. The value may be given as a percentage, and the abbreviation "v/v" for "volume per volume" may be used to indicate this choice. Conversion of solubility values Conversion between these various ways of measuring solubility may not be trivial, since it may require knowing the density of the solution — which is often not measured, and cannot be predicted. While the total mass is conserved by dissolution, the final volume may be different from both the volume of the solvent and the sum of the two volumes.[7] Moreover, many solids (such as acids and salts) will dissociate in non-trivial ways when dissolved; conversely, the solvent may form coordination complexes with the molecules or ions of the solute. In those cases, the sum of the moles of molecules of solute and solvent is not really the total moles of independent particles solution. To sidestep that problem, the solubility per mole of solution is usually computed and quoted as if the solute does not dissociate or form complexes -- that is, by pretending that the mole amount of solution is the sum of the mole amounts of the two substances. Qualifiers used to describe extent of solubility The extent of solubility ranges widely, from infinitely soluble (without limit, i. e. miscible[2]) such as ethanol in water, to essentially insoluble, such as titanium dioxide in water. A number of other descriptive terms are also used to qualify the extent of solubility for a given application. For example, U.S. Pharmacopoeia gives the following terms, according to the mass msv of solvent required to dissolve one unit of mass msu of solute:[8] (The solubilities of the examples are approximate, for water at 20-25 °C.) Term range Example g/dL msv/msu Very soluble 0) or exothermic (ΔH

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